MEASUREMENT OF D₂O IN WATER USING
²H(γ,n)¹H REACTION

M.F. Rahimi¹*, H. Peyrovan² and A. Izadpanah³

¹ Department of Physics, Faculty of Sciences, Ferdowsi University, Mashhad, Islamic Republic of Iran
² Department of Nuclear Physics, Atomic Energy Organization of Iran, Tehran, Islamic Republic of Iran
³ Department of Physics, Gorgan University, Gorgan, Islamic Republic of Iran

Abstract

Utilization of Photodisintegration reaction of Deuterium, ²H(γ,n)¹H, is a precise technique for determining the percentage of ²H in tritiated and natural water. The ²³Na(n,γ)²⁴Na reaction is a suitable γ-source for producing the ²H(γ,n)¹H reaction. Tehran Research Reactor was used as a neutron source for producing ²⁴Na isotopes. The BF₃ detectors were used to detect the neutron emission of H₂ content of water. We obtained a linear calibration plot by repeating the experiment with many samples of different concentrations of tritiated water. With this simple portable apparatus which have been used for this purpose for the first time, and corresponding linear plot, we can monitor and determine easily and accurately the ²H concentration of an unknown sample.

Keywords: Gamma source (Na²⁴); Tehran research reactor; Thermal neutron; BF₃ detectors; D₂O water; Linear plot

Introduction

Advancements in nuclear physics have provided powerful means for chemical analysis. Bombardment of selected nuclei with particles or gamma rays, to produce certain nuclear reaction, and detecting the emitted radiation is a well-established analytical process in nuclear technology. Thermal neutrons have been extensively used in neutron activation analysis. Although gamma irradiation did not find wide application, it has been utilized in Deuterium Photodisintegration [1-3].

When gamma rays of sufficient energy, (greater than the neutron binding energy) hit a light element, a probability exists that a neutron will be emitted by the energized nuclei. The majority of nuclei have a neutron binding energy of 5 MeV or more, except ⁷⁄₈ Be (1.67 MeV) or ¹⁄₂ H (2.23 MeV). Gamma rays from artificially produced radio-nuclei can serve in Photodisintegration reactions. A standard ²²⁸Th or ²⁴Na gamma source is placed at the center of a vessel filled with tritiated water (D₂O). The yield, the number of neutrons produced per second by 1 g of Be or D₂O at a distance of 1 cm derived from 1 curie of any of various gamma sources is given in Table 1.

The advantages of Na or Th sources are:

* E-mail: fa-rahimi@yahoo.com
(a) Reproducible neutron yield.
(b) The choice of $E_n$ by choosing a suitable value of $E_{\gamma}$.
(c) The wide variety of artificial gamma sources which are available from atomic reactors.
(d) The neutrons produced are more mono-energetic than those obtained from $(\alpha, n)$ sources.

Wattenberg [4,5] gives the equation of the emitted neutrons for the energy $(E_{\gamma}):$

$$E_{\gamma} = \frac{A-1}{A} \left[ E_{\gamma} - Q - \frac{E_{\gamma}^2}{1862(A-1)} \right] +$$

where, for $A(\gamma, n)A-1$ reactions:
- $A$: The mass of target nucleus
- $E_{\gamma}$: The energy of the incident gamma rays in MeV [the threshold energy, in MeV, for the $(\gamma, n)$ reaction] on target nucleus
- $Q$: 2.2246 MeV for $^2H(\gamma, n)^1H$
- $\theta$: The angle between the path of gamma rays and the M direction of the emitted neutrons.

When deuterium nuclei undergo Photodisintegration, they emit neutrons with relatively high energies. These neutrons must be slowed down to improve their detection efficiency. The availability of efficient neutron detectors and the electronics necessary to obtain a count will provide a nondestructive means for determining deuterium oxide in water.

$^2H(\gamma, n)^1H$ Reaction

Bombarding a water sample with a flux of gamma rays produces this reaction. Under the effect of these rays, the present deuterium ($^2H$ or D) nuclei will disintegrate to produce neutrons with an energy of about 0.2 MeV (Table 1). After being slowed down and detected, the monitored neutrons can be associated to the $^2H$ concentration in the water sample.

Safety considerations limit the intensity of gamma rays that can be utilized in this monitoring technique. The main disadvantage of Na$^{24}$ is that it has a short life and must be reactivated as needed.

**Neutron Detection**

Fast neutrons are of little interest to us, but after being moderated, the resulting thermal neutrons are very efficient to produce nuclear reactions. This not only provides a source of artificially produced isotopes, but also gives a means of detecting neutrons.

The monitored neutrons can be related to the $^2H_2O$ concentrations in water sample. An equation for $^2H_2O$ detection sensitivity is derived elsewhere [3]. This equation is expressed as:

$$S = \frac{R}{N_D G} = \frac{\sigma A \epsilon}{4 \pi r^2}$$

where
- $S$: Detection sensitivity
- $R$: Moderated neutron count rate (n/s)
- $N_D$: Number of Deuterium nuclei
- $G$: Gamma source strength ($\gamma$/s)
- $\sigma$: $^2H(\gamma, n)^1H$ Cross-section (cm$^2$)
- $A$: Correction factor for gamma ray attenuation
- $\epsilon$: Efficiency of neutron detector(s):

$$\epsilon = \epsilon(r) = \sum_{i=1}^{N} \epsilon(r_i),$$

where
- $\epsilon(r_i)$: The sum of the efficiencies for all neutron detectors.
- $r_i$: Source to Deuterium distance (cm). A few remarks concerning the variables that determine S are in order:
  - $\sigma$: The 2754-KeV gamma of Na$^{24}$ (15 h) has a $\sigma = 1.5 \times 10^{-27}$ cm$^2$. This $\sigma$ is 23% larger than that of the 2615-KeV gamma of $^{228}$Th (1.9 y), which has been used in some $^2H(\gamma, n)^1H$ monitors [3].
  - $A$: The attenuation for gamma rays with energies greater than 2000-KeV is close to unity for typical probe moderators [3] ($A \approx 1$ in Equation 1).
Table 1. Characteristics of some radioactive ($\gamma,n$) sources [4]

<table>
<thead>
<tr>
<th>$\gamma$-Source</th>
<th>Target</th>
<th>Half-life</th>
<th>$E_{\gamma}$ (MeV)</th>
<th>$E_n$ (MeV)</th>
<th>Standard yield</th>
</tr>
</thead>
<tbody>
<tr>
<td>$^{24}$Na</td>
<td>D$_2$O</td>
<td>14.8 h</td>
<td>2.76</td>
<td>0.22</td>
<td>$27 \times 10^4$</td>
</tr>
<tr>
<td>$^{24}$Na</td>
<td>Be</td>
<td>14.8 h</td>
<td>2.76</td>
<td>0.83</td>
<td>$13 \times 10^4$</td>
</tr>
<tr>
<td>$^{88}$Y</td>
<td>Be</td>
<td>87 d</td>
<td>1.9, 2.8</td>
<td>0.16</td>
<td>$10^5$</td>
</tr>
<tr>
<td>$^{88}$Y</td>
<td>D$_2$O</td>
<td>87 d</td>
<td>2.8</td>
<td>0.31</td>
<td>$0.3 \times 10^4$</td>
</tr>
<tr>
<td>$^{124}$Sb</td>
<td>Be</td>
<td>60 d</td>
<td>1.7</td>
<td>0.025</td>
<td>$19 \times 10^4$</td>
</tr>
<tr>
<td>Rd-Th*</td>
<td>D$_2$O</td>
<td>1.9 y</td>
<td>2.26</td>
<td>0.197 ± 0.010</td>
<td>$9.5 \times 10^4$</td>
</tr>
</tbody>
</table>

*Rd-Th source emits some neutrons through ($\alpha,n$) reactions with light elements around it, yielding an undesirable background.

Table 2. Basic radiations of $^{24}$Na and $^{228}$Th

<table>
<thead>
<tr>
<th>Nuclide</th>
<th>Radiations Produced by</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>$\alpha$</td>
</tr>
<tr>
<td>$^{24}$Na</td>
<td>4.17 max (0.003%)</td>
</tr>
<tr>
<td></td>
<td>1.389 max (100%)</td>
</tr>
<tr>
<td>$^{228}$Th</td>
<td>5.43 (71%)</td>
</tr>
<tr>
<td></td>
<td>5.34 (28%)</td>
</tr>
</tbody>
</table>

ε: Neutron detector efficiency depends on the neutron moderator and the geometry. Winn *et al.* used Helium ($^3$H) detectors in their work [3].

At low neutron energies, the total cross section for Boron (B) is higher than that of Hydrogen [6]. Therefore, the use of B detectors will probably improve the detection efficiency.

r$_g$: Minimizing Source to Deuterium distance, maximize neutron production for a given gamma intensity [3].

**Statistical Considerations**

Using statistical considerations, we can express the count rate limit $\Delta R$, corresponding to just-detected D$_2$O, as follows [3]:

$$\Delta R = 2\left(\frac{R_b}{t}\right)^2 + \frac{1}{t}$$

Determining the $R_b$ (Count/min) in a particular interval of time $t$ enables us to find $\Delta R$ according to Equation 2. For any particular detecting system, $S$ can be easily calculated. Hence, by rearranging Equation 1, the detection limit $\Delta N_D$ can be expressed by:

$$\Delta N_D = \frac{\Delta R}{SG}$$

For a 1- mCi source, and for various pipe geometries, the minimum detection limit ($\Delta N_D$)$_{min}$ varies 5 to about 100 $\mu$l/cm for D$_2$O streams. In other words, for the best pipe geometry ($\Delta N_D$)$_{min}$ is:

$$\Delta N_D = 5 \mu l \ D_2O = 1.5 \times 10^{20} \ D_2O \ Molecules$$

Since natural water contains about one D$_2$O molecule in every 7000 water molecules, then it is expressed that $7000 \times 0.005 = 35$ ml of natural water is required to avail $1.5 \times 10^{20}$ D$_2$O molecules. This volume is fairly small and does not present practical difficulties.

**Instrumentation**

All the instruments needed are usually available in any nuclear research center. This apparatus is portable, which is very important factor for monitoring any residual heavy water in pipes during the maintenance of power reactors.
**Design of Experiment**

The BF\textsubscript{3} detectors are surrounded by attenuators and the whole apparatus confined within a 37.0 × 37.0 × 30 cm\textsuperscript{3} cubic box. Inside of it, six BF\textsubscript{3} detectors are placed for collecting neutrons. The detectors are located on the circumference of a circle that has a 14 cm diameter. The center of the circle coincides with the center of the box and six BF\textsubscript{3} detectors are located along vertexes of a hexagon. The specific locations for the gamma source and the water sample are at the center of box. As the gamma source is from $^{24}$Na, it must be separated from the water sample. For this purpose, the source container and water sample are designed in the form of coaxial cylinder (Figs. 1 and 2). The outer

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**Figure 1.** Top view of apparatus.

**Figure 2.** Side view of apparatus.
cylinder is made of lead to prevent the gamma rays from interacting with D$_2$O. Solid paraffin is used for attenuating the fast neutrons produced from the $^3$H($\gamma$,n)$^4$H reactions.

**Source Preparation**

In preparing the gamma source, 0.7 grams of Na$_2$CO$_3$ (99.9 percent pure) were poured into a polyethylene capsule, measuring 5-mm in diameter and 28 mm in height. This capsule is placed inside a second container made of aluminum. This ensemble was irradiated for 4 min with a flux of $10^{12}$ n/sec cm$^2$ inside the core of the research reactor at the Atomic-Energy Organization of Iran (AEOI). The activity of the source, after interrupting bombardment, was 40.35 mCi.

**Electronic Circuit**

The electronic circuit used (Fig. 3), includes the probes and the electronic units. All of the BF$_3$ detectors in this experiment are connected to a preamplifier. The bias of BF$_3$-detectors operating voltage for detection system was 1250 V. A Multi Channel Analyzer (MCA) collects the amplified pulses. To eliminate the gamma rays from the background, the lower limit voltage of the SCA is adjusted to 3 V. This voltage is less than any typical pulse produced by neutrons (4.5 V), and more than any gamma rays from the background. Integrating the surface under a spectrum curve, at 400-second intervals does the count (Fig. 4). After determining the background, the experiment is repeated using 200 cm$^3$ water samples, containing different D$_2$O percentages. The heavy water (tritiated) samples used are 99.8% pure. For calibrating this system, the experiment was repeated with 21 different concentrations of heavy water. The concentrations ranged from 0.5% to 25.0%. We neglected the error in the D$_2$O concentration (this concentration is less than 150 ppm or 0.15% in ordinary water). The results, after adjusting for errors, are given in Figure 5. The corrections we have performed were background counts, coefficient reduction of source activity, and adjusting the statistics for errors in counting. The count rate during 400 sec, ranged from $280 \pm 20$ for 0.5% D$_2$O to $8050 \pm 90$ for 25% D$_2$O. A linear calibration plot was obtained. We tested this plot for a few determined mixtures. For example, using a 16% of D$_2$O, we should have a count rate equal to 4860. The experimental result gave 4760, which is in good agreement with the calibrated plot.
Conclusions

(1) Utilization of the simple reaction $^2\text{H}(\gamma,n)^1\text{H}$, can provide a very powerful analytical technique for determination of Deuterium in water. The technique is so efficient that it can even detect $\text{D}_2\text{O}$ in depleted water.

(2) For preliminary experiments, if $^{24}\text{Na}$ (15 h) is not obtainable, one can utilize $^{228}\text{Th}$ (1.9 y) as the gamma source. As noted previously, the latter nuclide can provide more-lasting gamma source that is necessary for any practical works.

(3) Parameters like the volume of sample, the detector type, and different Gamma source, examined in order to reach the most efficient detection capacity. The recommended BF$_3$-detectors will possibly serve the purpose.

References